

Reactions Continued...

- Energy in reactions:

- Exothermic \rightarrow produce thermal energy, causing a temperature rise
- Endothermic \rightarrow require a lot of energy, so temperature drops

- Catalyst

- Every reaction happens at some rate.
- A catalyst speeds up the rate of reaction.
- The catalyst is never changed and is never used up.
- Each catalyst has one specific job.

Types of Reactions:

- Addition (Synthesis)

- Two or more substances combine to form a new compound



- Decomposition:

- Compound breaks into two or more smaller compounds



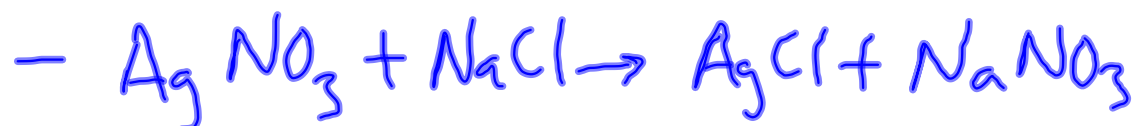
- Single Displacement:

- One element replaces another in a compound



- Double Displacement:

- Ions switch places (same charge)



- Acid-Base Reaction:

- Double displacement that has products of a salt and H_2O



- Combustion:

- Carbon compounds combine with oxygen to release carbon dioxide and water (and energy)

